Section 6.10

Types of chemical reactions: single and double displacement

Single displacement:

Are chemical changes that involve an element and a compound as reactants. One element displaces or replaces another element from a compound. The general formula is:

A + BC 🡪 B + AC

Example:

Magnesium + silver nitrate 🡪 silver + magnesium nitrate

Mg + AgNo3 🡪 Ag +Mg(NO3)2

Bromine + Calcium iodide 🡪 iodine + calcium bromide

Br2 + CaI2 🡪 I2 + CaBr2

Double displacement

Occurs when elements in different coupmounds displace or exchenge places. General formula:

AB + XY 🡪 AY + XB

Example:

Lead(II)nitrate + Potasium iodide 🡪 Lead(II)iodide + potassium nitrate

Pb(NO3)2 + KI 🡪 PbI2 + KNO3