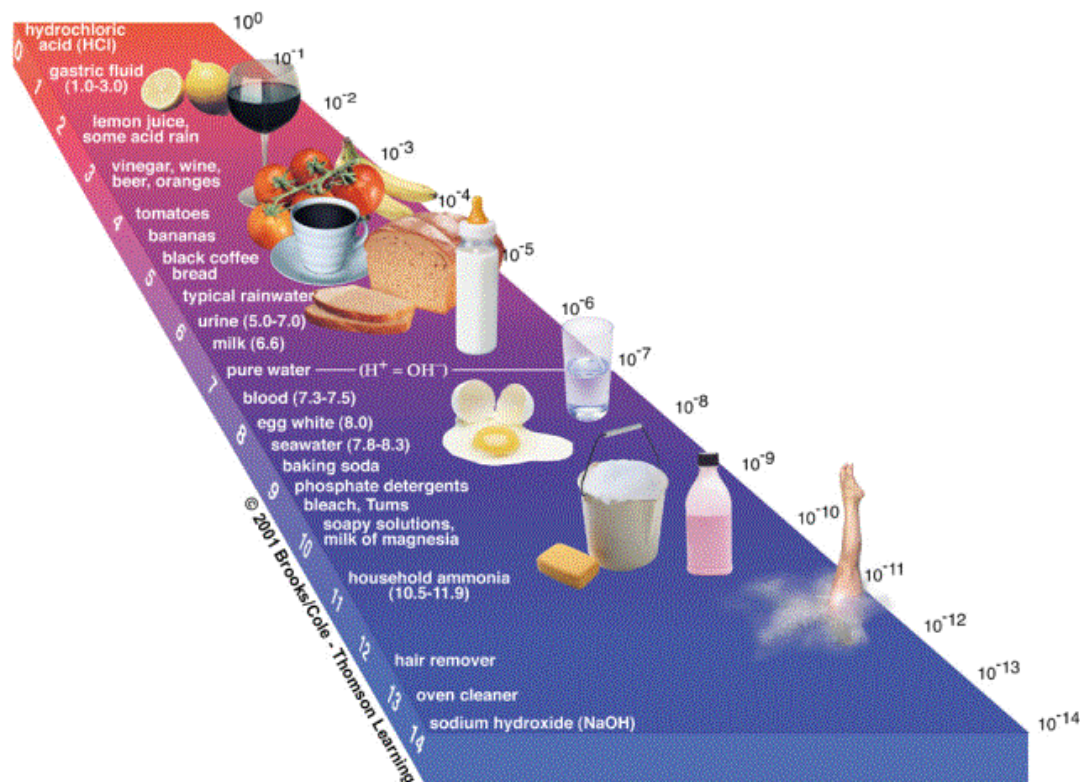


8.2

Review naming compounds

http://www.zerobio.com/drag_gr9/molecule_magic/magic.htm

Properties of acids and bases



Acids are:

1. Ionic compounds
2. Sour tasting
3. Water soluble
4. Reactive and can combine with many products
5. Good conductors of electricity
6. Break apart to form hydrogen ions

Chart of common acids

Vinegar	$\text{HC}_2\text{H}_3\text{O}_2$	Salad dressing
Citric acid	$\text{HC}_6\text{H}_7\text{O}_6$	oranges lemons
Ascorbic acid	$\text{HC}_6\text{H}_7\text{O}_6$	vitamin C
Lactic acid	$\text{HC}_3\text{H}_5\text{O}_3$	sour milk
Carbonic acid	H_2CO_3	carbonated drinks
Acetylsalicylic acid	$\text{HC}_9\text{H}_7\text{O}_4$	Aspirin
Sulfuric acid	H_2SO_4	car batteries

Bases are:

1. Ionic compounds
2. Bitter tasting
3. Water soluble
4. Form hydroxide ions

Common bases:

Sodium hydroxide	NaOH	drain cleaner
Potassium Hydroxide	KOH	soap cosmetics
Aluminum hydroxide	$\text{Al}(\text{OH})_3$	antacids
Ammonium Hydroxide	NH_4OH	ammonia wood cleaner
Sodium bicarbonate	NaHCO_3	baking soda
Potassium sulfite	K_2SO_3	food preserve

Distinguishing between acids and bases with chemical names

Acids have hydrogen H at the beginning

Bases have hydroxide OH at the end, some contain carbonate or bicarbonate and react with water to form hydroxide ions.

<http://www.youtube.com/watch?v=RF40cI2O16U&feature=related>

Read pages 293 - 295

Answer:

1. Describe three chemical reactions that involve acids.
2. Write a word equation to represent each reaction
3. Identify each of the following substances as acid or base
 - a. Potassium hydroxide
 - b. HClO_3
 - c. $\text{Mg}(\text{OH})_2$
 - d. HNO_3
 - e. Potassium bicarbonate
4. Read the included article and answer the questions on acid indigestion.
 - a. http://www.execulink.com/~ekimmel/secretin_0.htm
 - b. Read article stomach imbalances.

